Chemistry

Extra Credit

Mole Conversions

1. A sample contains 0.564 moles of FeO. How many grams of FeO is this?

2. For # above, how many formula units does this represent? How many atoms of oxygen are present?

3. 3.How many formula units are in 85g of $AgNO_3$?

Empirical Formulas and Percent Composition

1. What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen?

2. If the molar mass of the compound in problem 1 is 110 grams/mole, what's the molecular formula?

Write the molecular formula of the following compound

3. A compound with an empirical formula of C2OH4 and a molar mass of 88 grams per mole.

Answer the following questions

4. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound?